Chapter 7: Kinetic molecular theory explains the characteristics of solids, liquids, and gases.

Key Ideas:

◼ The particle model of matter describes how particles of solids are closer together than particles of liquids. Particles of a gas are spread far apart. (7.1)

◼ Kinetic molecular theory describes how adding energy to particles makes them move faster and farther apart. (7.1)

◼ Adding or removing energy from matter can cause changes in the state of matter. (7.1)

◼ Liquids and gases are fluids, forms of matter that can flow. (7.2)

◼ Density is a way to describe how closely particles are packed together in a solid, liquid,   
or gas. (7.2)

◼ Density is calculated by dividing mass by volume. (7.2)

Chapter 7 Key Terms:

States of Matter

* matter
* mass
* volume
* kinetic energy
* kinetic molecular theory

Matter and Temperature

* thermal expansion
* thermal contraction
* thermal energy
* heat
* temperature

Changes in State

* melting
* evaporation
* condensation
* solidification
* sublimation
* deposition
* melting point
* boiling point

Fluids and Density

* density
* displacement
* fluid